

Name:

Finding Energy for reactions that have a energy value given. (beside or within the equation)

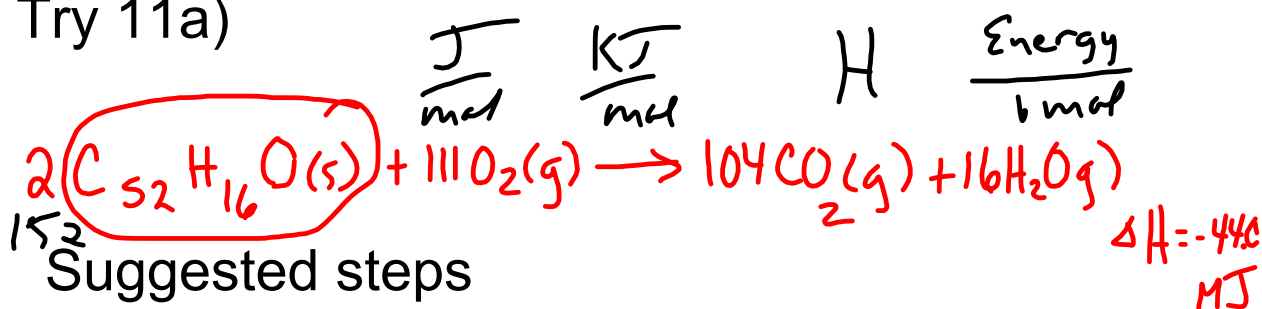
Enthalpy-heat content of a system

$$\Delta H = nH$$

Questions like this: p 516 14 and 15
p 535 58

Practise sheet

Try 11a)



Suggested steps

1. Find energy per mol from the given reaction. (This is molar enthalpy) $\Delta H = \frac{440 MJ}{2 mol} = 220 MJ/mol$
(whatever the question is referring to)

2. Find mol of substance the question is referring to. $n = \frac{100000 g}{656.68 g/mol} = 152 mol$

3. Use $\Delta H = nH$

$$\Delta H = 152 mol \times 22 \frac{MJ}{mol} = 3.35 \times 10^3 MJ$$

$3.35 \times 10^6 kJ$